I. Multiple Choice Questions (Type-I)

1. Addition of water to alkynes occurs in acidic medium and in the presence of Hg$^{2+}$ ions as a catalyst. Which of the following products will be formed on addition of water to but-1-yne under these conditions.

   (i) CH$_3$-CH$_2$-CH$_2$-C-H
   (ii) CH$_3$-CH$_2$-C-CH$_3$

   (iii) CH$_3$-CH$_2$-C-OH + CO$_2$
   (iv) CH$_3$-C-OH + H-C-H

2. Which of the following compounds is most reactive towards nucleophilic addition reactions?

   (i) CH$_3$-C-H
   (ii) CH$_3$-C-CH$_3$

   (iii) C=H
   (iv) C=CH$_3$

3. The correct order of increasing acidic strength is__________.

   (i) Phenol < Ethanol < Chloroacetic acid < Acetic acid
   (ii) Ethanol < Phenol < Chloroacetic acid < Acetic acid
   (iii) Ethanol < Phenol < Acetic acid < Chloroacetic acid
   (iv) Chloroacetic acid < Acetic acid < Phenol < Ethanol
4. Compound Ph—O—C—Ph can be prepared by the reaction of
   (i) Phenol and benzoic acid in the presence of NaOH
   (ii) Phenol and benzoyl chloride in the presence of pyridine
   (iii) Phenol and benzoyl chloride in the presence of ZnCl₂
   (iv) Phenol and benzaldehyde in the presence of palladium

5. The reagent which does not react with both, acetone and benzaldehyde.
   (i) Sodium hydrogensulphite
   (ii) Phenyl hydrazine
   (iii) Fehling's solution
   (iv) Grignard reagent

6. Cannizaro's reaction is not given by __________.
   (i) \[
   \begin{array}{c}
   \text{CHO} \\
   \text{CH₃}
   \end{array}
   \]
   (ii) \[
   \begin{array}{c}
   \text{CHO} \\
   \text{C}\text{H₃}
   \end{array}
   \]
   (iii) \[
   \begin{array}{c}
   \text{H} \\
   \text{CHO}
   \end{array}
   \]
   (iv) \[
   \begin{array}{c}
   \text{CH₃CHO}
   \end{array}
   \]

7. Which product is formed when the compound \[
   \begin{array}{c}
   \text{CHO}
   \end{array}
   \]
   is treated with concentrated aqueous KOH solution?
   (i) \[
   \begin{array}{c}
   \text{C} \\
   \text{O} \\
   \text{K}^- \\
   \text{O} \\
   \text{K}^+
   \end{array}
   \]
   (ii) \[
   \begin{array}{c}
   \text{O} \\
   \text{C} \\
   \text{K}^- \\
   \text{O} \\
   \text{O} \\
   \text{K}^+
   \end{array}
   \]
   (iii) \[
   \begin{array}{c}
   \text{O} \\
   \text{C} \\
   \text{O} \\
   \text{K}^- \\
   \text{O} \\
   \text{K}^+
   \end{array}
   \]
   (iv) \[
   \begin{array}{c}
   \text{O} \\
   \text{C} \\
   \text{O} \\
   \text{K}^- \\
   \text{O} \\
   \text{K}^+
   \end{array}
   \]
8. \[ \text{CH}_3\text{C} = \text{CH} \xrightarrow{40\% \text{H}_2\text{SO}_4, 1\% \text{HgSO}_4} \text{Isomerisation} \xrightarrow{} \text{CH}_3\text{C} - \text{CH}_3 \]

Structure of \('\text{A}'\) and type of isomerism in the above reaction are respectively.

(i) Prop-1-en-2-ol, metamerism
(ii) Prop-1-en-1-ol, tautomerism
(iii) Prop-2-en-2-ol, geometrical isomerism
(iv) Prop-1-en-2-ol, tautomerism

9. Compounds \(\text{A}\) and \(\text{C}\) in the following reaction are ________.

\[
\begin{align*}
\text{CH}_3\text{CHO} & \xrightarrow{[\text{i}]} \text{CH}_3\text{MgBr} & \xrightarrow{[\text{ii}]} \text{H}_2\text{O} & \xrightarrow{\text{H}_2\text{SO}_4, \Delta} \text{CH}_3\text{SO}_3\text{H} & \xrightarrow{\text{oxidation}} \text{C} \\
& \text{[A]} & \text{(B)} & \text{(C)} \\
\end{align*}
\]

(i) identical
(ii) positional isomers
(iii) functional isomers
(iv) optical isomers

10. Which is the most suitable reagent for the following conversion?

\[
\begin{align*}
\text{CH}_3\text{C} = \text{CH} - \text{CH} - \text{CH}_3 & \xrightarrow{} \text{CH}_3\text{C} = \text{CH} - \text{CH} - \text{CH}_3 \\
& \xrightarrow{} \text{CH}_3\text{C} = \text{CH} - \text{CH}_3 - \text{COH} \\
\end{align*}
\]

(i) Tollens’ reagent
(ii) Benzoyl peroxide
(iii) \(\text{I}_2\) and \(\text{NaOH}\) solution
(iv) \(\text{Sn}\) and \(\text{NaOH}\) solution

11. Which of the following compounds will give butanone on oxidation with alkaline \(\text{KMnO}_4\) solution?

(i) Butan-1-ol
(ii) Butan-2-ol
(iii) Both of these
(iv) None of these

12. In Clemmensen Reduction carbonyl compound is treated with ________.

(i) Zinc amalgam + \(\text{HCl}\)
(ii) Sodium amalgam + \(\text{HCl}\)
(iii) Zinc amalgam + nitric acid
(iv) Sodium amalgam + \(\text{HNO}_3\)
II. Multiple Choice Questions (Type-II)

Note: In the following questions two or more options may be correct.

13. Which of the following compounds do not undergo aldol condensation?

(i) \( \text{CH}_3-\text{CHO} \)  
(ii) \( \text{CH}_3-\text{C}-\text{CHO} \)

(iii) \( \text{CH}_3-C-\text{CH}_3 \)  
(iv) \( \text{CH}_3-C-\text{CHO} \)

14. Treatment of compound \( \text{Ph}-\text{O}-\text{C}-\text{Ph} \) with NaOH solution yields

(i) Phenol  
(ii) Sodium phenoxide  
(iii) Sodium benzoate  
(iv) Benzophenone

15. Which of the following conversions can be carried out by Clemmensen Reduction?

(i) Benzaldehyde into benzyl alcohol  
(ii) Cyclohexanone into cyclohexane  
(iii) Benzoyl chloride into benzaldehyde  
(iv) Benzophenone into diphenyl methane

16. Through which of the following reactions number of carbon atoms can be increased in the chain?

(i) Grignard reaction  
(ii) Cannizaro's reaction  
(iii) Aldol condensation  
(iv) HVZ reaction

17. Benzophenone can be obtained by ___________.

(i) Benzoyl chloride + Benzene + AlCl_3  
(ii) Benzoyl chloride + Diphenyl cadmium  
(iii) Benzoyl chloride + Phenyl magnesium chloride  
(iv) Benzene + Carbon monoxide + ZnCl_2

171. Aldehydes, Ketones and Carboxylic Acids
18. Which of the following is the correct representation for intermediate of nucleophilic addition reaction to the given carbonyl compound [A]:

\[ \begin{array}{c}
\text{Nu} \\
\text{(i)} \\
\text{Nu} \\
\text{Nu} \\
\text{(ii)} \\
\text{(iii)} \\
\text{Nu} \\
\text{(iv)} \\
\end{array} \]

III. Short Answer Type

19. Why is there a large difference in the boiling points of butanal and butan-1-ol?

20. Write a test to differentiate between pentan-2-one and pentan-3-one.

21. Give the IUPAC names of the following compounds:

\[ \begin{array}{c}
\text{(i)} \\
\text{(ii)} \\
\text{(iii)} \\
\text{(iv)} \\
\end{array} \]

22. Give the structure of the following compounds.
(i) 4-Nitropropiophenone
(ii) 2-Hydroxycyclopentanecarbaldehyde
(iii) Phenyl acetaldehyde

23. Write IUPAC names of the following structures.

\[ \begin{array}{c}
\text{(i)} \\
\text{(ii)} \\
\text{(iii)} \\
\end{array} \]

24. Benzaldehyde can be obtained from benzal chloride. Write reactions for obtaining benzal chloride and then benzaldehyde from it.
25. Name the electrophile produced in the reaction of benzene with benzoyl chloride in the presence of anhydrous AlCl₃. Name the reaction also.

26. Oxidation of ketones involves carbon-carbon bond cleavage. Name the products formed on oxidation of 2, 5-dimethylhexan-3-one.

27. Arrange the following in decreasing order of their acidic strength and give reason for your answer.
   CH₃CH₂OH, CH₃COOH, CICH₂COOH, FCH₂COOH, C₆H₅CH₂COOH

28. What product will be formed on reaction of propanal with 2-methylpropanal in the presence of NaOH? What products will be formed? Write the name of the reaction also.

29. Compound 'A' was prepared by oxidation of compound 'B' with alkaline KMnO₄. Compound 'A' on reduction with lithium aluminium hydride gets converted back to compound 'B'. When compound 'A' is heated with compound B in the presence of H₂SO₄ it produces fruity smell of compound C to which family the compounds 'A', 'B' and 'C' belong to?

30. Arrange the following in decreasing order of their acidic strength. Give explanation for the arrangement.
   C₆H₅COOH, FCH₂COOH, NO₂CH₂COOH

31. Alkenes (C=C) and carbonyl compounds (C=O), both contain a π bond but alkenes show electrophilic addition reactions whereas carbonyl compounds show nucleophilic addition reactions. Explain.

32. Carboxylic acids contain carbonyl group but do not show the nucleophilic addition reaction like aldehydes or ketones. Why?

33. Identify the compounds A, B and C in the following reaction.
   \[
   \text{CH₂Br} \xrightarrow{\text{Mg/ether}} (\text{A}) \xrightarrow{(i)} \text{CO₂} \xrightarrow{(ii) \text{Water}} (\text{B}) \xrightarrow{\Delta} (\text{C})
   \]

34. Why are carboxylic acids more acidic than alcohols or phenols although all of them have hydrogen atom attached to a oxygen atom [O—H]? 

35. Complete the following reaction sequence.
   \[
   \text{CH₂—C—CH₃} \xrightarrow{(i)} \text{CH₃MgBr} \xrightarrow{\text{Na metal}} (\text{A}) \xrightarrow{\text{H₂O}} (\text{B}) \xrightarrow{\text{C₆H₅Br}} (\text{C})
   \]

36. Ethylbenzene is generally prepared by acetylation of benzene followed by reduction and not by direct alkylation. Think of a possible reason.

37. Can Gatterman-Koch reaction be considered similar to Friedel Craft's acylation? Discuss.
IV. Matching Type

Note: Match the items of Column I and Column II in the following questions.

38. Match the common names given in Column I with the IUPAC names given in Column II.

<table>
<thead>
<tr>
<th>Column I (Common names)</th>
<th>Column II (IUPAC names)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(i) Cinnamaldehyde</td>
<td>(a) Pentanal</td>
</tr>
<tr>
<td>(ii) Acetophenone</td>
<td>(b) Prop-2-enal</td>
</tr>
<tr>
<td>(iii) Valeraldehyde</td>
<td>(c) 4-Methylpent-3-en-2-one</td>
</tr>
<tr>
<td>(iv) Acrolein</td>
<td>(d) 3-Phenylprop-2-enal</td>
</tr>
<tr>
<td>(v) Mesityl oxide</td>
<td>(e) 1-Phenylethanone</td>
</tr>
</tbody>
</table>

39. Match the acids given in Column I with their correct IUPAC names given in Column II.

<table>
<thead>
<tr>
<th>Column I (Acids)</th>
<th>Column II (IUPAC names)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(i) Phthalic acid</td>
<td>(a) Hexane-1,6-dioic acid</td>
</tr>
<tr>
<td>(ii) Oxalic acid</td>
<td>(b) Benzene-1,2-dicarboxylic acid</td>
</tr>
<tr>
<td>(iii) Succinic acid</td>
<td>(c) Pentane-1,5-dioic acid</td>
</tr>
<tr>
<td>(iv) Adipic acid</td>
<td>(d) Butane-1,4-dioic acid</td>
</tr>
<tr>
<td>(v) Glutaric acid</td>
<td>(e) Ethane-1,2-dioic acid</td>
</tr>
</tbody>
</table>

40. Match the reactions given in Column I with the suitable reagents given in Column II.

<table>
<thead>
<tr>
<th>Column I (Reactions)</th>
<th>Column II (Reagents)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(i) Benzophenone → Diphenylmethane</td>
<td>(a) LiAlH₄</td>
</tr>
<tr>
<td>(ii) Benzaldehyde → 1-Phenylethanol</td>
<td>(b) DIBAL—H</td>
</tr>
<tr>
<td>(iii) Cyclohexanone → Cyclohexanol</td>
<td>(c) Zn(Hg)/Conc. HCl</td>
</tr>
<tr>
<td>(iv) Phenyl benzoate → Benzaldehyde</td>
<td>(d) CH₃MgBr</td>
</tr>
</tbody>
</table>

41. Match the example given in Column I with the name of the reaction in Column II.

<table>
<thead>
<tr>
<th>Column I (Example)</th>
<th>Column II (Reaction)</th>
</tr>
</thead>
<tbody>
<tr>
<td>( \text{CH}_3\text{--C--Cl} + \text{H}_2 \xrightarrow{\text{Pd-C/BaSO}_4} \text{CH}_3\text{--C--H} )</td>
<td>(a) Friedel Crafts acylation</td>
</tr>
</tbody>
</table>
V. Assertion and Reason Type

Note: In the following questions a statement of assertion followed by a statement of reason is given. Choose the correct answer out of the following choices.

(i) Assertion and reason both are correct and reason is correct explanation of assertion.
(ii) Assertion and reason both are wrong statements.
(iii) Assertion is correct statement but reason is wrong statement.
(iv) Assertion is wrong statement but reason is correct statement.
(v) Assertion and reason both are correct statements but reason is not correct explanation of assertion.

42. Assertion: Formaldehyde is a planar molecule.
   Reason: It contains sp² hybridised carbon atom.

43. Assertion: Compounds containing —CHO group are easily oxidised to corresponding carboxylic acids.
   Reason: Carboxylic acids can be reduced to alcohols by treatment with LiAlH₄.

175. Aldehydes, Ketones and Carboxylic Acids
44. **Assertion**: The α-hydrogen atom in carbonyl compounds is less acidic.
   **Reason**: The anion formed after the loss of α-hydrogen atom is resonance stabilised.

45. **Assertion**: Aromatic aldehydes and formaldehyde undergo Cannizzaro reaction.
   **Reason**: Aromatic aldehydes are almost as reactive as formaldehyde.

46. **Assertion**: Aldehydes and ketones, both react with Tollens’s reagent to form silver mirror.
   **Reason**: Both, aldehydes and ketones contain a carbonyl group.

**VI. Long Answer Type**

47. An alkene 'A' (Mol. formula C\textsubscript{4}H\textsubscript{10}) on ozonolysis gives a mixture of two compounds 'B' and 'C'. Compound 'B' gives positive Fehling's test and also forms iodoform on treatment with I\textsubscript{2} and NaOH. Compound 'C' does not give Fehling's test but forms iodoform. Identify the compounds A, B and C. Write the reaction for ozonolysis and formation of iodoform from B and C.

48. An aromatic compound 'A' (Molecular formula C\textsubscript{9}H\textsubscript{8}O) gives positive 2, 4-DNP test. It gives a yellow precipitate of compound 'B' on treatment with iodine and sodium hydroxide solution. Compound 'A' does not give Tollens's or Fehling's test. On drastic oxidation with potassium permanganate it forms a carboxylic acid 'C' (Molecular formula C\textsubscript{4}H\textsubscript{4}O\textsubscript{2}), which is also formed along with the yellow compound in the above reaction. Identify A, B and C and write all the reactions involved.

49. Write down functional isomers of a carbonyl compound with molecular formula C\textsubscript{4}H\textsubscript{4}O. Which isomer will react faster with HCN and why? Explain the mechanism of the reaction also. Will the reaction lead to the completion with the conversion of whole reactant into product at reaction conditions? If a strong acid is added to the reaction mixture what will be the effect on concentration of the product and why?

50. When liquid 'A' is treated with a freshly prepared ammoniacal silver nitrate solution, it gives bright silver mirror. The liquid forms a white crystalline solid on treatment with sodium hydrogensulphite. Liquid 'B' also forms a white crystalline solid with sodium hydrogensulphite but it does not give test with ammoniacal silver nitrate. Which of the two liquids is aldehyde? Write the chemical equations of these reactions also.
ANSWERS

I. Multiple Choice Questions (Type-I)
1. (ii)  2. (i)  3. (iii)  4. (ii)  5. (iii)  6. (iv)

II. Multiple Choice Questions (Type-II)
13. (ii), (iv); [Hint : in compounds (ii) and (iv) α-hydrogen is absent.]
14. (i), (iii) 15. (ii), (iv) 16. (i), (iii) 17. (i), (ii) 18. (i), (ii)

III. Short Answer Type
19. [Hint : Butan-1-ol has higher boiling point due to intermolecular hydrogen bonding.]
20. [Hint : Iodoform test]
21. (i) 3-Phenylprop-2-enal  (ii) Cyclohexanecarbaldehyde
   (iii) 3-oxopentanal  (iv) But-2-enal
22. (i) \[\text{NO}_2\] \[\begin{array}{c}C \quad \text{CH}_3 \end{array}\]
   (ii) \[\begin{array}{c}C \quad \text{OH} \end{array}\]
   (iii) \[\begin{array}{c}C \quad \text{CHO} \end{array}\]
23. (i) Ethane-1, 2-dial  (ii) Benzene-1, 4-dicarbaldehyde
   (iii) 3-Bromobenzaldehyde
24. See NCERT textbook for Class XII
25. \(\text{C}_6\text{H}_5\text{CO}^+\) benzoyliumcation or \(\text{C}_6\text{H}_5\text{C}^+\text{Cl}^-\text{AlCl}_3\); Friedel Craft’s acylation reaction.
26.

\[
\begin{array}{c}
\text{CH}_3 \quad \text{O} \\
\text{CH}_3 \\
\text{CH}_3 \\
\text{CH}_3 \\
\text{CH}_3 \quad \text{[O]} \\
\text{CH}_3 \\
\text{CH}_3 \\
\end{array}
\quad \text{CH}_3 \quad \text{CH} \quad \text{CH}_2 \quad \text{CH}_2 \quad \text{CH}_3 \quad \text{CH}_3 \\
\text{CH}_3 \quad \text{CH} \quad \text{CH}_2 \quad \text{CH} \quad \text{CH}_3 \quad \text{COOH} \\
\text{CH}_3 \quad \text{CH} \quad \text{CH}_2 \quad \text{CH}_2 \quad \text{CH}_3 \quad \text{COOH} \\
\text{C}_6\text{H}_5\text{COOH} + \text{HCOOH} \quad \left(\text{2-Methylpropanoic acid}\right) \quad \left(\text{3-Methylbutanoic acid}\right)
\end{array}
\]

(Propan-2-one)

177. Aldehydes, Ketones and Carboxylic Acids
27. **Hint**: $\text{FCH}_2\text{COOH} > \text{ClCH}_2\text{COOH} > \text{C}_6\text{H}_5\text{CH}_2\text{COOH} > \text{CH}_3\text{COOH} > \text{CH}_3\text{CH}_2\text{OH}$

28. It is cross Aldol condensation

\[
\begin{align*}
\text{CH}_3\text{CH}_2\text{CHO} + \text{CH}_3\text{CHCH}_3 & \rightarrow \text{CH}_3\text{CH}_2\text{CH} - \text{C} - \text{CHO} + \text{CH}_3\text{CH} - \text{CH} = \text{C} - \text{CHO} \\
& \quad + \text{CH}_3\text{CH} - \text{C} - \text{CHO} + \text{CH} - \text{CH} = \text{C} - \text{CHO}
\end{align*}
\]

29. 'A' is a carboxylic acid, 'B' is an alcohol and 'C' is an ester.

30. $\text{NO}_2\text{CH}_2\text{COOH} > \text{FCH}_2\text{COOH} > \text{CH}_3\text{COOH}$

**Hint**: electron withdrawing effect.

31. **Hint**: Carbon atom in carbonyl compounds acquires slight positive charge and is attacked by nucleophile.

32. **Hint**: Due to resonance as shown below the partial positive charge on carbonyl carbon atom is reduced.

\[
\begin{align*}
\text{O} & \quad \text{C} _\text{=} \quad \text{H} \\
\text{C} & \quad \text{O} \quad \text{C} _\text{=} \quad \text{O} \quad \text{H}
\end{align*}
\]

33. $\text{A} = \text{CH}_3\text{MgBr} \quad \text{B} = \text{CH}_3\text{COOH} \quad \text{C} = \text{CH}_3\text{C} - \text{O} - \text{CH}_3$

34. **Hint**: Compare the stability of anion formed after the loss of $\text{H}^+$ ion. More stable the anion formed, more easy will be the dissociation of $\text{O} - \text{H}$ bond, stronger will be the acids.

35. **Hint**: $\text{A} = \text{CH}_3\text{O} - \text{CH}_3 \quad \text{B} = \text{CH}_3\text{C} - \text{CH}_3 \quad \text{C} = \text{CH}_3\text{C} - \text{O} - \text{CH}_3$

**IV. Matching Type**

38. (i) — (d), (ii) — (c), (iii) — (a), (iv) — (b), (v) — (c)

39. (i) — (b), (ii) — (e), (iii) — (d), (iv) — (a), (v) — (c)

40. (i) — (c), (ii) — (d), (iii) — (a), (iv) — (b)

41. (i) — (c), (ii) — (d), (iii) — (a), (iv) — (b), (v) — (f), (vi) — (c)
V. Assertion and Reason Type

42. (i) 43. (v) 44. (iv) 45. (iii) 46. (iv)

VI. Long Answer Type

47. \[ \text{CH}_3-\text{CH}=\text{C}-\text{CH}_3 + \text{O}_2 \xrightarrow{\text{[i]} \text{Zn}/\text{H}_2\text{O}} \text{H}_2\text{C}-\text{CHO} + \text{O}=\text{C}-\text{CH}_3 \]

2-Methylbut-2-ene  \hspace{1cm} (A) \hspace{1cm} (B) \hspace{1cm} (C)

Other isomers of ‘A’ will not give products corresponding to the given test.

48. **Hint:**

49. \[ \text{CH}_3\text{CH}_2\text{CHO} \hspace{1cm} \text{CH}_3\text{COCH}_3 \]

- Compound I will react faster with HCN due to less steric hinderance and electronic reasons than II.
- No, it is a reversible reaction. Hence equilibrium is established.
- Addition of acid inhibits the reaction because the formation of NCN ions is prevented.

50. **Hint:** Liquid ‘A’