

**CBSE**  
**Class XII Chemistry**  
**Board Paper – 2018**

**Time: 3 Hours**

**Total Marks: 70**

**General Instructions:**

- Please check this question paper contains 26 questions.  
15 min time has been allotted to read this question paper.
- The question paper has been distributed at 10.15 am.  
From 10.15 am to 10.30 am, the students will read the question paper only and will not write any answer on answer book during this period.

1. Write the coordination number and oxidation state of Platinum in the complex  $[\text{Pt}(\text{en})_2\text{Cl}_2]$ . [1]
2. Analysis shows that FeO has a non-stoichiometric composition with formula  $\text{Fe}_{0.95}\text{O}$ . Give reason. [1]
3. Out of chlorobenzene and benzyl chloride, which one gets easily hydrolysed by aqueous NaOH and why? [1]
4. Write the IUPAC name of the following: [1]  
$$\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3 - \text{C} - \text{CH} - \text{CH}_3 \\ | \quad | \\ \text{C}_2\text{H}_5 \quad \text{OH} \end{array}$$
5.  $\text{CO}_{(g)}$  and  $\text{H}_2_{(g)}$  react to give different products in the presence of different catalysts. Which ability of the catalyst is shown by these reactions? [1]
6. Among the hydrides of Group-15 elements, which have the [2]  
(a) Lowest boiling point?  
(b) Maximum basic character?  
(c) Highest bond angle?  
(d) Maximum reducing character?
7. Calculate the freezing point of a solution containing 60 g of glucose (Molar mass =  $180\text{g mol}^{-1}$ ) in 250 g of water. ( $K_f$  of water =  $1.86\text{ K kg mol}^{-1}$ ) [2]

8. How do you convert the following? [2]  
(a) Ethanal to Propanone  
(b) Toluene to Benzoic acid

OR

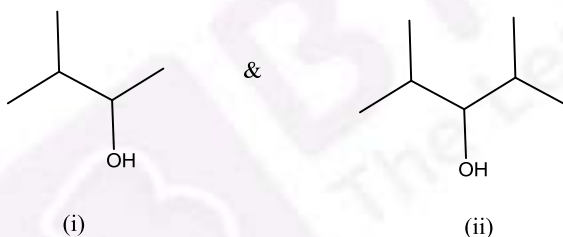
Account for the following:

- (a) Aromatic carboxylic acids do not undergo Friedel-Crafts reaction.  
(b)  $pK_a$  value of 4-nitrobenzoic acid is lower than that of benzoic acid.

9. Complete and balance the following chemical equations: [2]  
(a)  $Fe^{2+} + MnO_4^- + H^+ \rightarrow$   
(b)  $MnO_4^- + H_2O + I^- \rightarrow$

10. For the reaction [2]  
 $2N_2O_5(g) \rightarrow 4NO_2(g) + O_2(g)$   
The rate of formation of  $NO_2(g)$  is  $2.8 \times 10^{-3} M s^{-1}$ .  
Calculate the rate of disappearance of  $N_2O_5(g)$ .

11. [3]  
(a) Identify the chiral molecule in the following pair:



- (b) Write the structure of the product when chlorobenzene is treated with methyl chloride in the presence of sodium metal and dry ether.
- (c) Write the structure of the alkene formed by dehydrohalogenation of 1-bromo-1-methylcyclohexane with alcoholic KOH.
12. A first order reaction is 50% completed in 40 minutes at 300 K and in 20 minutes at 320 K. Calculate the activation energy of the reaction. (Given:  $\log 2 = 0.3010$ ,  $\log 4 = 0.6021$ ,  $R = 8.314 JK^{-1} mol^{-1}$ ) [3]
13. An element X' (At. mass a 40  $g mol^{-1}$ ) having f.c.c. structure, has unit cell edge length of 400 pm. Calculate the density of 'X' and the number of unit cells in 4 g of 'X'. ( $N_A = 6.022 \times 10^{23} mol^{-1}$ ) [3]

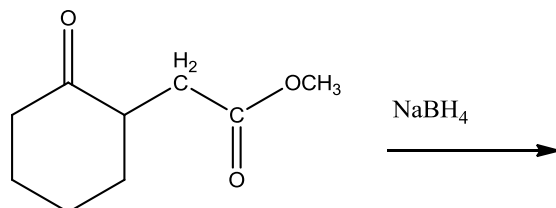
- 14.** Give reasons for the following: [3]
- (a) Measurement of osmotic pressure method is preferred for the determination of molar masses of macromolecules such as proteins and polymers.
  - (b) Aquatic animals are more comfortable in cold water than in warm water.
  - (c) Elevation of boiling point of 1 M KCl solution is nearly double than that of 1 M sugar solution.
- 15.** What happens when [3]
- (a) a freshly prepared precipitate of  $\text{Fe}(\text{OH})_3$  is shaken with a small amount of  $\text{FeCl}_3$  solution?
  - (b) persistent dialysis of a colloidal solution is carried out?
  - (c) an emulsion is centrifuged?
- 16.** Write the chemical reactions involved in the process of extraction of Gold. Explain the role of dilute NaCN and Zn in this process. [3]
- 17.** (A), (B) and (C) are three non-cyclic functional isomers of a carbonyl compound with molecular formula  $\text{C}_4\text{H}_8\text{O}$ . Isomers (A) and (C) give positive Tollens' test whereas isomer (B) does not give Tollens test but gives positive Iodoform test. Isomers (A) and (B) on reduction with  $\text{Zn}(\text{Hg})/\text{conc. HCl}$  give the same product (D). [3]
- (a) Write the structures of (A), (B), (C) and (D).
  - (b) Out of (A), (B) and (C) isomers, which one is least reactive towards addition of HCN?
- 18.** [3]
- (a) Why is bithional added to soap?
  - (b) What is tincture of iodine? Write its one use.
  - (c) Among the following, which one acts as a food preservative? Aspartame, Aspirin, Sodium Benzoate, Paracetamol
- 19.** [3]
- (a) Write the formula of the following coordination compound : iron(III) hexacyanoferrate(II)
  - (b) What type of isomerism is exhibited by the complex  $(\text{Co}(\text{NH}_3)_5\text{Cl})\text{SO}_4$ ?
  - (c) Write the hybridisation and number of unpaired electrons in the complex  $[\text{CoF}_6]^{3-}$ . (Atomic No. of Co = 27)
- 20.** Define the following with an example of each: [3]
- (a) Polysaccharides
  - (b) Denatured protein
  - (c) Essential amino acids

**OR**

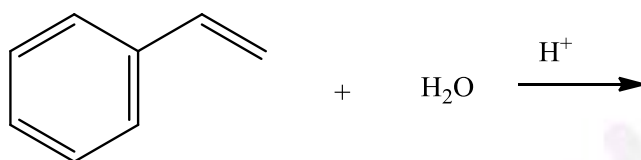
- (a) Write the product when D-glucose reacts with con.  $\text{HNO}_3$ .  
 (b) Amino acids show amphoteric behaviour. Why?  
 (c) Write one difference between  $\alpha$ -helix and  $\beta$ -pleated structures of proteins.

21. Write the structures of the main products in the following reactions: [3]

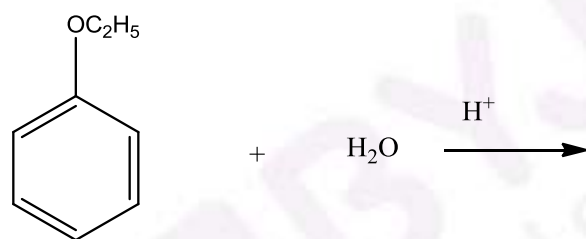
(i)



(ii)



(iii)



22. Give reasons [3]

- (a)  $E^0$  value for  $\text{Mn}^{3+}/\text{Mn}^{2+}$  couple is much more positive than that for  $\text{Fe}^{3+}/\text{Fe}^{2+}$ .  
 (b) Iron has higher enthalpy of atomization than that of copper.  
 (c)  $\text{Sc}^{3+}$  is colourless in aqueous solution whereas  $\text{Ti}^{3+}$  is coloured.

23. Shyam went to a grocery shop to purchase some food items. The shopkeeper packed all the items in polythene bags and gave them to Shyam. But Shyam refused to accept the polythene bags and asked the shopkeeper to pack the items in paper bags. He informed the shopkeeper about the heavy penalty imposed by the government for using polythene bags. The shopkeeper promised that he would use paper bags in future in place of polythene bags. [4]

Answer the following:

- (a) Write the values (at least two) shown by Shyam.  
 (b) Write one structural difference between low-density polythene and high-density polythene.  
 (c) Why did Shyam refuse to accept the items in polythene bags?  
 (d) What is a biodegradable polymer? Give an example.

24.

[5]

(a) Write the reactions involved in the following:

- (i) Hofmann bromamide degradation reaction
- (ii) Diazotisation
- (iii) Gabriel phthalimide synthesis

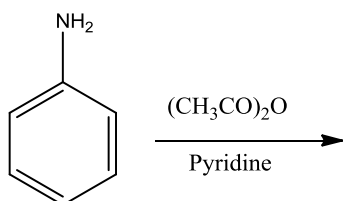
(b) Give reasons

- (i)  $(\text{CH}_3)_2\text{NH}$  is more basic than  $(\text{CH}_2)_3\text{N}$  in an aqueous solution.
- (ii) Aromatic diazonium Salts are more stable than aliphatic diazonium salts.

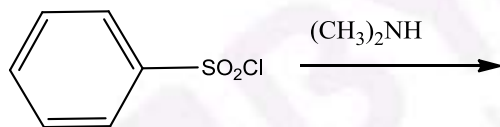
**OR**

(a) Write the annums of the main products of the following reactions: [5]

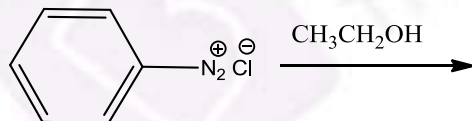
(i)



(ii)

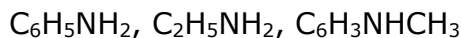


(iii)



(b) Give a simple chemical test to distinguish between Aniline and N,N-dimethylaniline.

(c) Arrange the following in the increasing order of their  $\text{p}K_b$  values:



25.

[5]

(a) Give reasons:

- (i)  $\text{H}_3\text{PO}_3$  undergoes disproportionation reaction but  $\text{H}_3\text{PO}_4$  does not.
- (ii) When  $\text{Cl}_2$  reacts with excess of  $\text{F}_2$ ,  $\text{ClF}_3$  is formed and not  $\text{FCl}_3$ .
- (iii) Dioxygen is a gas while Sulphur is a solid at room temperature.

(b) Draw the structures of the following:

- (i)  $\text{XeF}_4$
- (ii)  $\text{HClO}_3$

**OR**

- (a) When concentrated sulphuric acid was added to an unknown salt present in a test tube a brown gas (A) was evolved. This gas intensified when copper turnings were added to this test tube. On cooling, the gas (A) changed into a colourless solid (B). [5]
- Identify (A) and (B).
  - Write the structures of (A) and (B).
  - Why does gas (A) change to solid on cooling?

- (b) Arrange the following in the decreasing order of their reducing character:  
HF, HCl, HBr, HI

- (c) Complete the following reaction:  
 $\text{XeF}_4 + \text{SbF}_5 \rightarrow$

**26.** [5]

- (a) Write the cell reaction and calculate the e.m.f. of the following cell at 298 K:  
 $\text{Sn (a)} | \text{Sn}^{2+} (0.004 \text{ M}) || \text{H}^+ (0.020 \text{ M}) | \text{H}_2(\text{g}) (1 \text{ bar}) | \text{Pt (s)}$  (Given :  $E^\circ_{\text{Sn}^{2+}/\text{Sn}} = -0.14 \text{ V}$ )
- (b) Give reasons:
- On the basis of  $E^\circ$  values,  $\text{O}_2$  gas should be liberated at anode but it is  $\text{Cl}_2$  gas which is liberated in the electrolysis of aqueous NaCl.
  - Conductivity of  $\text{CH}_3\text{COOH}$  decreases on dilution.

**OR**

- (a) For the reaction



$$\Delta G^\circ = -43600 \text{ J at } 25^\circ\text{C}.$$

Calculate the e.m.f. of the cell.

$$[\log 10^{-n} = -n]$$

- (b) Define fuel cell and write its two advantages.