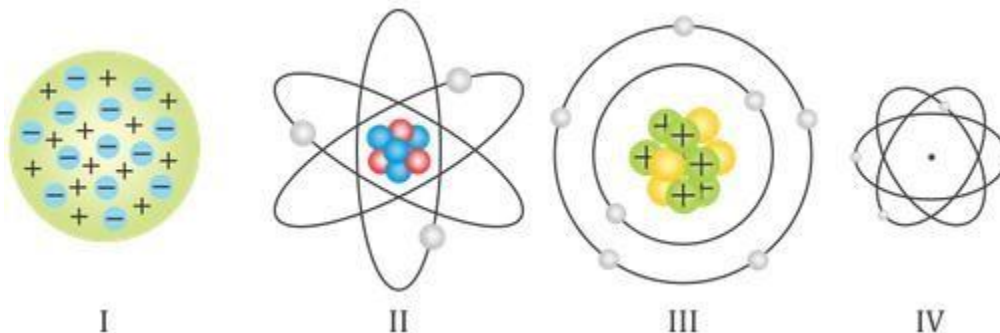


ICSE Class 8 Chemistry Important Questions

1. Which of the following pictures represent the Bohr's model of an atom?



- (a) I
- (b) II
- (c) III
- (d) IV

2. Sodium forms sodium cation by loss of?

- (a) One electron
- (b) Two electrons
- (c) Three electrons
- (d) Four electrons

3. Hydrogen frees metals from their

- (a) Sulphates
- (b) Oxides
- (c) Nitrates
- (d) Chlorides

4. State the electronic configuration of the following atoms:

- i. Atom 'A' (Atomic number = 12)
- ii. Atom 'B' (Atomic number = 19)
- iii. Atom 'C' (Atomic number = 7)
- iv. Atom 'D' (Atomic number = 26)
- v. Atom 'E' (Atomic number = 10)

5. Fill in the blanks and rewrite the sentences:

- i. When few drops of phenolphthalein are added to sodium hydroxide, the solution turns_____.
- ii. Separation of components of air such as liquid nitrogen and liquid oxygen is possible by_____.
- iii. The three states of matter are classified on the basis of differences in certain _____ properties.
- iv. A mixture which has different composition and properties in different parts of their

- v. mass is called a _____ mixture.
_____ is used in the laboratory for the preparation of hydrogen using dilute hydrochloric acid.
6. State whether the following statements are true or false.
Rewrite the false statement.
- A compound is a pure substance composed only of one kind of element combined chemically in a fixed proportion by mass.
 - A chemical reaction in which two or more substances combine to form a single product is called a combination reaction or synthesis.
 - In the formation of ammonia from hydrogen and nitrogen iron is used as positive catalyst.
 - Magnesium has atomic number 12 and atomic mass number 24.
 - The insoluble solid settles down in a beaker is called as sediment.
7. Name the following:
- Conversion of a solid to a liquid on heating
 - Conversion of a liquid to a vapour (or gas)
 - Conversion of a vapour (or gas) to a liquid
 - Conversion of a liquid to a solid
 - The outermost shell or orbit of an atom.
8. Give distinguishing explanation between alpha rays, beta rays and gamma rays.
9. What is a metal reactivity series? What are its important features?
10. Define the following:
- Lamp black
 - Sugar charcoal
 - Chemical equation
 - Orbit
 - Bone charcoal
11. Match the name of the radical given in **Column A** with its formula given in **Column B**.

Sr. No.	Column A (Name of the radical)	Column B (Formula)
1.	Chlorate	HCO_3^-
2.	Bicarbonate	MnO_4^-
3.	Bisulphate	Cu^{2+}
4.	Permanganate	ClO_3^-
5.	Cupric	HSO_4^-

12. Write the formula of the given compounds:

- i. Acetic acid
- ii. Sodium hydroxide
- iii. Sulphuric acid
- iv. Hydrochloric acid
- v. Ammonium hydroxide

13. Write the electronic configuration of the following elements:

- i. Sodium
- ii. Chlorine
- iii. Hydrogen
- iv. Nitrogen
- v. Oxygen

14. Differentiate between compound and mixture.

15. What is atomicity? Give one example each of mono atomic and diatomic molecule.

16. Give the chemical equations for a reaction of potassium with

(a) Oxygen

(b) Water

17. Write the main features of Rutherford's atomic model.

18. How can you prove that hydrogen burns in air to produce water?

19. Differentiate between oxidation reaction and reduction reaction.

20. What are protons, neutrons and electrons?