

**MP BOARD CLASS 12 CHEMISTRY
PREVIOUS YEAR QUESTION -2016**

E-444 (H/E)

CHEMISTRY 2016

Time : 3 Hours]

Class : 12th

[M. M. : 75

Instructions-

- (i) Attempt all the questions.
- (ii) Question Nos. 1 to 4 are objective types. Carries total 20 marks.
- (iii) Question Nos. 5 to 8, each question carries 2 marks. (Word limit 30 words)
- (iv) Question Nos. 9 to 13, each question carries 4 marks. (Word Limit 75 words)
- (v) Question Nos. 14 to 16, each question carries 5 marks (Word Limit 120 words)
- (vi) Question Nos. 17 to 18, each question carries 6 marks (Word Limit 150 words)
- (vii) Internal choice is given to question Nos. 5 to 18.

Q. 1. Choose the correct option: 5 × 1 = 5

- (a) Which is not found in R.N.A.

(i) Thymine	(ii) Ureacel
(iii) Adinine	(iv) Guwanine
- (b) Which Noble gas does not have Octer complete:

(i) Helium	(ii) Neon
(iii) Argon	(iv) Krypton
- (c) Which Halogen sublimates:

(i) Chlorine	(ii) Bromine
(iii) Iodine	(iv) Fluorine
- (d) For Increasing of electro-conductivity in a solid crystal, mixing of impurities is known as:

(i) Schottky defect	(ii) Frenkel defect
(iii) Doping	(iv) Electronic-Defect
- (e) Formula of unit cell Density is:

(i) $\frac{ZM}{a^3 N_0}$	(ii) $\frac{ZN_0}{a^3 M}$
(iii) $\frac{N_0 a^3}{MZ}$	(iv) $\frac{Z}{MN_0}$

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- Q. 2.** Fill in the blanks: 5 × 1 = 5
- Co-ordination number of Sodium in NaCl is
 - Substance which are attracted in Magnetic field are called
 - The Potential value of standard hydrogen electrode is
 - Alkaline solution of HgCl_2 and KI is called
 - Main product of mustard oil reaction is
- Q. 3.** Write answer in one word of each: 5 × 1 = 5
- What is the name of reaction for preparation of methyl isocyanide.
 - Conversion of precipitate into colloidal solution is known as.
 - What is the name of Disaccharides Sugar present in milk.
 - Which types of Isomerism are in $[\text{Co}(\text{NH}_3)_5\text{Br}]\text{SO}_4$ and $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Br}$.
 - What is the name of reactions which are initiated by the radiations.
- Q. 4.** Match the pairs correctly: 5 × 1 = 5
- | Column 'A' | Column 'B' |
|-----------------------|--|
| (a) Radon | (i) Ligand |
| (b) Hinsberg Reagent | (ii) Clotting of Blood |
| (c) Catalyst Promoter | (iii) $\text{C}_6\text{H}_5\text{SO}_2\text{Cl}$ |
| (d) E.D.T.A | (iv) Molybdenum |
| (e) Vitamin K | (v) Inert gas |
- Q. 5.** What do you understand by denaturation of Protein. 2
- (OR)** Give two difference between D.N.A. and R.N.A.
- Q. 6.** Give two difference between double salt and complex salt. 2
- (OR)** Name is IUPAC System:
- $\text{K}_2[\text{PtCl}_6]$
 - $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$
- Q. 7.** What is tyndall effect. 2
- (OR)** Give two differences between Lyophilic colloids and Lyophobic colloids.
- Q. 8.** What are interhalogen compounds. 2
- (OR)** Write down name, structural formula and Hybridization of two Xenon compounds.
- Q. 9.** Write short notes on: 4
- Threshold Energy
 - Energy of Activation
- (OR)** What do you understand by order of reaction? Give three examples of first order of reactions.
- Q. 10.** Write four alloys of Aluminium with its names, composition and uses. 4

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- (OR) Write names, formula and uses of following compounds:
- Hematite
 - Silver Glans
 - Luner-Caustic
 - Corrosive Sublimate
- Q. 11. How will you make (obtain) from chloroform: 4
- Chloretone
 - Phenyl Isocynide
 - Acetylene
 - Salicaldehyde
- (OR) What happens when: (Write equations)
- Acetone is heated with Alkaline solution of Iodine.
 - Chlorobenzene is heated with chloral in presence of conc- H_2SO_4 .
 - Chlorobenzene is heated with Sodium in presence of ether.
 - Ethyl bromide is heated with Alcoholic-KOH.
- Q. 12. Difference between primary, secondary and tertiary alcohol through Victor Meyer's method only Equations. 4
- (OR) (i) How will you obtain from phenol:
- 2, 4, 6 Tribromo Phenol
 - Picric-Acid
- (ii) What is the reaction of Diethyl ether with HI-Acid.
- Q. 13. Describe the preparation of acetic acid through quick-vinegar process in following points: 4
- Diagram
 - Method
 - Equation of preparation
 - Any one precaution.
- (OR) What happens when:
- Acetal chloride is reduced with hydrogen in presence of Barium sulphate associated with paladium.
 - Benzaldehyde boils with 45% NaOH.
 - Ammonia reacts with Formaldehyde.
 - Acetic-Acid reacts with- PCl_5 .
- Q. 14. 10.07 gram Silver obtained during the circulations of electric of 5 ampere upto 30 minutes in silver nitrate Pot (cell). Find out electro-chemical equivalent of silver. If chemical equivalent of hydrogen is 0.00001036 then what will be equivalent weight of silver. 5
- (OR) (i) What do you understand by corrosion.
- (ii) Write electro-chemical theory of corrosion (Rust).

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- (iii) Write prevention (two) of corrosion.
- Q. 15. (a) Bleaching of flowers by chlorine is stable but bleaching of SO_2 is unstable. Why? (2.5 × 2 = 5)
- (b) Water is liquid while hydrogen sulphide is a gas at normal temperature. Why?
- (OR) Draw a Labelled diagram of Ostwald method for preparation of Nitric Acid and explain its preparation only by equation.
- Q. 16. (i) Give two names of Artificial sweeters. (1 + 2 + 2 = 5)
- (ii) Give definition of antibiotic and one example.
- (iii) Give definition of antihistamine drug with name and uses.
- (OR) (a) Write short notes on Sushrut.
- (b) Give the names of active ingredients and their one uses of following medicinal plants.
- (i) Amla
- (ii) Haldi
- (iii) Tulsi
- Q. 17. What do you understand by elevation of boiling point and molal elevation constant, with the help of this constant how can you find the molecular mass of non volatile solute? 6
- (OR) (a) Give two differences between Ideal and non-Ideal solutions.
- (b) Find out osmotic pressure of Glucose solutions of 5% sol. at 25°C temp while molecular mass of Glucose 180 and $R = 0.0821$ Litre Atmosphere.
- Q. 18. (a) Why Transition elements become paramagnetic? (3 + 3 = 6)
- (b) Explain oxidization properties of potassium permagnate in acidic medium. (any four only)
- (OR) (a) What are Lanthanide elemnts?
- (b) What do you understand by Lanthanide contraction?
- (c) Write chromyl chloride test with equation.