General instructions:

Read the following instructions very carefully and strictly follow them:

i. This question paper comprises four sections – A, B, C and D. There are 37 questions in the question paper. All questions are compulsory.

ii. Section A – Questions no. 1 to 20 are very short answer type questions, carrying 1 mark each. Answer these questions in one word or one sentence.

iii. Section B – Questions no. 21 to 27 are short answer type questions, carrying 2 marks each.

iv. Section C – Questions no. 28 to 34 are long answer type – I questions, carrying 3 marks each.

v. Section D – Questions no. 35 to 37 are long answer type – II questions, carrying 5 marks each.

vi. There is no overall choice in the question paper. However, an internal choice has been provided in 2 questions of two marks, 2 questions of three marks and all the 3 questions of five marks. You have to attempt only one of the choices in such questions.

vii. In addition to this, separate instructions are given with each section and question, wherever necessary.

viii. Use of calculators and log tables is not permitted.

SECTION – A

Read the given passage and answer the question number 1 to 5 that follow: 1 × 5 = 5

Organic compounds containing amine as functional group are present in a vivid variety of compounds, namely amino acids, hormones, neurotransmitters, DNA, alkaloids, dyes, etc. Drugs including nicotine, morphine, codeine and heroin, etc. which have physiological effects on humans also contain amino group in one form or another. Amines are basic because of the presence of lone pair of electrons on nitrogen. Addition of nitrogen into an organic framework leads to the formation of two families of molecules, namely amines and amides. As chemistry students, we must appreciate the versatility of nitrogen.

1. What are amino acids?
2. Why are amino acids amphoteric?
3. Give one point of difference between acidic and basic amino acids.
4. What are essential amino acids?
5. Name the linkage formed when carboxyl end of one amino acid condenses with amino end of other amino acid.

Question number 6 to 10 are one word answers: 1 × 5 = 5
6. Name the process where the metal is converted into a volatile compound and is collected elsewhere.

7. An organic compound is adsorbed on the surface of silica gel. Name the process of removing the organic compound from silica gel.

8. Give an example of a metal which can be purified by the process of distillation.

9. What type of isomerism is shown by the complex $\text{[Co(NH}_3\text{)}_5\text{NO}_2\text{]}\text{Cl}_2$?

10. Calculate the overall order of the reaction whose rate law expression was predicted as: 
    \[
    \text{Rate} = k[\text{NO}]^{3/2}[\text{O}]^{1/2}
    \]

Questions number 11 to 15 are multiple choice questions:  
\[1 \times 5 = 5\]

11. The correct IUPAC name of 

   (A) tert-butyl alcohol  
   (B) 2, 2-Dimethyl propanol  
   (C) 2-Methyl butan-2-ol  
   (D) 3-Methylbutan-3-ol

12. 50 mL of an aqueous solution of glucose $\text{C}_6\text{H}_{12}\text{O}_6$ (Molar mass 180 g/mol) $6.02 \times 10^{22}$ molecules. The concentration of the solution will be

   (A) 0.1 M  
   (B) 0.2 M  
   (C) 1.0 M  
   (D) 2.0 M

13. If the standard electrode potential of an electrode is greater than zero, then we can infer that its

   (A) reduced form is more stable compared to hydrogen gas.  
   (B) oxidised form is more stable compared to hydrogen gas.  
   (C) reduced and oxidised forms are equally stable.  
   (D) reduced form is less stable than the hydrogen gas.

14. Total number of unpaired electrons present in $\text{Co}^{3+}$ (Atomic number = 27) is

   (A) 2  
   (B) 7  
   (C) 3  
   (D) 5

15. The incorrect statement about interstitial compounds is

   (A) They are chemically reactive  
   (B) They are very hard.  
   (C) They retain metallic conductivity  
   (D) They have high melting point.
For question number 16 to 20, two statements are given - one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (i), (ii), (iii) and (iv) as given below:

(i) Both assertion (A) and reason (R) are correct statements, and reason (R) is the correct explanation of the assertion (A).

(ii) Both assertion (A) and reason (R) are correct statements, but reason (R) is not the correct explanation of the assertion (A).

(iii) Assertion (A) is correct, but reason (R) is incorrect statement.

(iv) Assertion (A) is incorrect, but reason (R) is correct statement.

16. Assertion (A): o-nitrophenol is a weaker acid than p-nitrophenol.
   Reason (R): Intramolecular hydrogen bonding makes ortho isomer weaker than para isomer.

17. Assertion (A): Albumin is a globular protein.
   Reason (R): Polypeptide chain coils around to give a straight chain.

18. Assertion (A): Boiling points of alkyl halides decrease in the order R-I > R-Br > R-Cl > R-F.
   Reason (R): Van der Waals forces decrease with increase in the size of halogen atom.

19. Assertion (A): Low spin tetrahedral complexes are rarely observed.
   Reason (R): The orbital splitting energies are not sufficiently large to forcing pairing.

20. Assertion (A): Bakelite is a thermosetting polymer.
   Reason (R): On heating, polymeric chain becomes a long and straight chain.

SECTION – B

21. Give reasons for the following:

   1 × 2 = 2

   (a) Bond angle in alcohol is slightly less than the tetrahedral angle.

   (b) C - OH bond length in CH₃OH is slightly more than the C - OH bond length in phenol.

22. Give one point of difference between the following:

   1 × 2 = 2

   (a) Tranquilizers and Analgesics
   (b) Antiseptics and Disinfectants

   OR

   Differentiate on the basis of chemical composition between cationic and anionic detergents. Also give one example of each category.

23. Complete and balance the following chemical equations:

   2
(a) \[ \text{MnO}_4^- + H_2O + I^- \rightarrow \]

(b) \[ \text{MnO}_4^2- + H^+ + I^- \rightarrow \]

24. A reaction is first order w.r.t. reactant A as well as w.r.t. reactant B. Give the rate law. Also give one point of difference between average rate and instantaneous rate.

25. Define adsorption isotherm. Give the empirical relationship between the quantity of gas adsorbed by unit mass of solid absorbent and pressure at a particular temperature.

OR

Define shape-selective catalysis. Name the process by which alcohols convert directly into gasoline and give a variety of hydrocarbons.

26. For an electrochemical cell:

\[ \text{Cu}^{2+}(aq) + \text{Ni}(s) \rightarrow \text{Ni}^{2+}(aq) + \text{Cu}(s), \]

Give the cell representation. Also write the Nernst equation for the above cell at 25°C.

27. Predict the state of the solute in the solution in the following situations:

(a) When ‘i’ is found to be more than one.
(b) When ‘i’ is found to be less than one.

SECTION – C

28. Give the structures of the monomers of the following polymers: 1 × 3 = 3

(a) Neoprene
(b) Nylon – 6, 6
(c) Dacron

OR

Write the names of monomers of the following polymers: 1 × 3 = 3

29. (a) Give the IUPAC name and electronic configuration of central metal atom in terms of t_{2g} and e_{g} of K_4[Mn(CN)_6].
(b) What is meant by ‘Chelate effect’? Give an example.  

[OR]

Write the hybridization and magnetic characters of the following complexes:  

(i) $[\text{Fe(CN)}_6]^{4-}$  (ii) $[\text{CoF}_6]^{3-}$  (iii) $[\text{Ni(CO)}_4]$  

[Atomic number : Fe = 26, Co = 27, Ni = 28]

30. Justify and arrange the following compounds namely  

\[
\begin{array}{c}
\text{CH}_3 \\
\text{CH}_2\text{Br} \\
\text{CH}_3
\end{array}
\]

in increasing order of reactivity towards the asked displacement namely:  

(a) $S_N1$  (b) $S_N2$

31. Account for the following:  

(a) Aniline is a weaker base compared to ethanamine.

(b) Aniline does not undergo Friedel Crafts reaction.

(c) Only aliphatic primary amines can be prepared by Gabriel Phthalimide synthesis.

32. Give the chemical reactions involved in the leaching of alumina from bauxite.

33. State Kohlrausch's law. Calculate the molar conductance of $\text{Sr(NO}_3\text{)}_2$. The molar ionic conductance of $\text{Sr}^{2+}$ and $\text{NO}_3^-$ ions are 119 S cm$^2$ mol$^{-1}$ and 728 cm$^2$ mol$^{-1}$ respectively.

34. An antifreeze solution is prepared by dissolving 31 g of ethylene glycol (Molar mass = 62 g mol$^{-1}$) in 600 g of water. Calculate the freezing point of the solution ($K_f$ for water = 1.86 K kg mol$^{-1}$)

35. (a) An organic compound ‘A’ having molecular formula C$_5$H$_{10}$O gives negative Tollens' test, forms n-pentane on Clemmensen reduction but doesn’t give iodoform test. Identify ‘A’ and give all the reactions involved.

(b) Carry out the following conversions:

(i) Propanoic acid to 2-Bromopropanoic acid

(ii) Benzoyl chloride to benzaldehyde

(c) How will you distinguish between benzaldehyde and acetaldehyde?

OR

(a) Complete the following sequence of reactions:

\[
\frac{1}{2} \times 4 = 2
\]
(i) Identify (A) to (D)

(ii) Give the IUPAC name of (A).

(b) How can you distinguish between:

(i) Ethanol and Propanone, and

(ii) Benzoic acid and Phenol?

36. (a) Visha plotted a graph between concentration of R and time for a reaction $R \rightarrow P$. On the basis of this graph, answer the following questions:

(i) Predict the order of reaction.

(ii) What does the slope of the line indicate?

(iii) What are the units of rate constant?

(b) A first order reaction takes 25 minutes for 25% decomposition. Calculate $t_{1/2}$.

[Given: $\log 2 = 0.3010$, $\log 3 = 0.4771$, $\log 4 = 0.6021$]
(a) The rate constant for a first order reaction is 60 s\(^{-1}\). How much time will it take to reduce the initial concentration of the reactant to its \(\frac{1}{16}\)th value? 3

(b) Write two factors that affect the rate of a chemical reaction. 1

(c) Write two conditions for the collisions to be effective collisions. 1

37. An amorphous solid ‘A’ which has a crown shaped structure, burns in air to form a gas ‘B’ which turns lime water milky. ‘B’ is also produced by roasting of sulphide ores. ‘B’ undergoes oxidation in the presence of V\(_2\)O\(_5\) to give ‘C’ and to carry out this oxidation low temperature and high pressure is mandatory to get a good yield of ‘C’. ‘C’ is then absorbed in H\(_2\)SO\(_4\) to give ‘D’. ‘D’ is then diluted to give a very important compound ‘E’. ‘E’ is largely responsible for the manufacture of variety of compounds in industry. ‘E’ in concentrated form, when combined with Cu metal, gives compound ‘F’.

From this description

(a) Elucidate the structure of ‘A’ to ‘F’.
(b) Give a balanced chemical equation for the conversion of ‘E’ to ‘F’.
(c) Give two important functions of ‘E’ in the chemical industry.

[OR]

(a) Give reasons for the following observations:
(i) Halogens are strong oxidising agents.
(ii) Noble gases have very low boiling points.
(iii) O and Cl have nearly same electronegativity, yet oxygen forms H bond while Cl doesn’t.

(b) Complete and balance the following chemical equations:
(i) \( \text{NaOH} + \text{Cl}_2 \rightarrow \)
(cold + dil.)
(ii) \( I^- (aq) + H_2O(l) + O_2(g) \rightarrow \)