

**CBSE Class 10 Science Question Paper Solution 2020 Set
31/4/1**

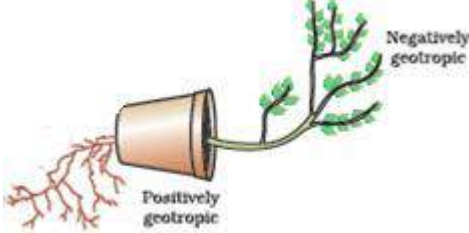
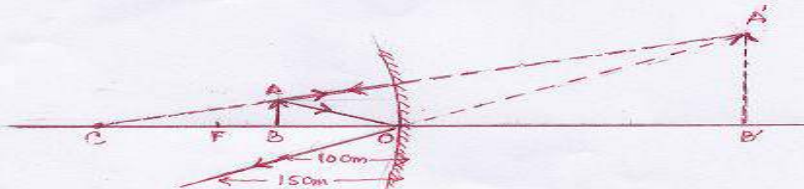
Series –JBB/4

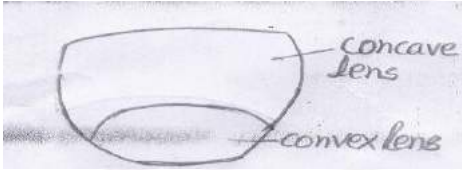
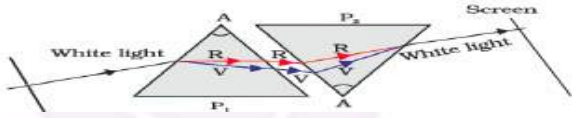
Set -1

Paper Code : 31/4/1

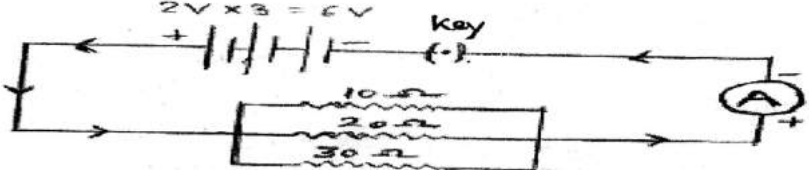
MARKING SCHEME –CLASS X SCIENCE (2019-20)			
QUESTION PAPER CODE : SET 31/4/1			
S.NO	VALUE POINTS/EXPECTED ANSWER	MARKS	TOTAL MARKS
SECTION A			
1.	5 valence electrons	1	1
2	The electric current generated /induced in a conductor by changing magnetic field around it.	1	1
3.	(a) The properties of elements are the periodic functions of their atomic masses. (b) To fill with undiscovered elements. (c) (ii)/RH ₄ , RO ₂ (d) (i)/Atoms of an element with similar chemical properties but different atomic masses.	1 1 1 1	4
4.	(a) Use of separate bins for plastic and paper ; separation of biodegradable and non biodegradable waste or any other. (b) <ul style="list-style-type: none"> • Packaging of articles like water, food, milk, biscuits etc. • Disposable utility items –bowls, tumblers, plates , leaves etc. (c) By providing cloth /jute /earthen pots and utensils/ paper or any other material for the similar purposes. (d) <ul style="list-style-type: none"> • Yes • The action of microbes is tested in the laboratory creating the same conditions as in the landfill. 	½+ ½ ½ + ½ 1 ½ ½	4
5.	(C) / Valves ensure that the blood does not flow backwards.	1	1
6.	(A)/ takes place in yeast during fermentation. OR (A)/ small intestine	1	1
7.	(B) /fusion of nuclei of male and female gamete.	1	1
8.	(A) /1 Ω OR (B) /half	1 1	1
9.	(C) /direction of the induced current.	1	1
10.	(B)/ The nucleus of Uranium is bombarded with high energy neutrons. OR (A)/ Biomass	1	1
11.	(C) /various interlinked food chains in an ecosystem.	1	1
12.	Note: Treat all answers as correct. Give full credit even if not attempted.	1	1
13.	(b) /Both (A) and (R) are true but (R) is not the correct explanation of the assertion (A).	1	1
14.	(a) /Both (A) and (R) are true and (R) is the correct explanation of the assertion (A).	1	1
SECTION B			
15.	(a) Lead iodide; Yellow colour (b) Pb(NO ₃) ₂ + 2 KI → PbI ₂ + 2KNO ₃ Note : ½ marks to be deducted if reaction is not balanced.	½ + ½ 1	

	<p>(c) Double displacement reaction ; Due to exchange of ions. OR</p> <ul style="list-style-type: none"> Fats and oils become rancid. Observable changes – <ol style="list-style-type: none"> Change in taste Change in smell Three ways of prevention :- <ol style="list-style-type: none"> Addition of antioxidants / substance which prevent oxidation. Keeping food in air tight containers. Use of nitrogen gas in packaged food. 	$\frac{1}{2} + \frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$	3																								
16.	<table border="1" style="width: 100%;"> <thead> <tr> <th style="width: 50%; text-align: center;">Galvanisation</th> <th style="width: 50%; text-align: center;">Alloying</th> </tr> </thead> <tbody> <tr> <td>1. Coating a layer of zinc metal on the metal.</td> <td>1. Mixing of a metal with metal or non-metal.</td> </tr> <tr> <td>2. Not a homogeneous mixture.</td> <td>2. Homogeneous mixture.</td> </tr> <tr> <td>3. No change in physical properties of metals takes place.</td> <td>3. Change in physical properties of metals takes place.</td> </tr> <tr> <td>4. The process is an outcome of the reactivity of metals.</td> <td>4. Reactivity of metals do not play any role in it.</td> </tr> <tr> <td>5. Prevents rusting only.</td> <td>5. Some alloys may prevent rusting and also used for other advantages.</td> </tr> </tbody> </table> <p style="text-align: right;">(Any Three)</p> <p style="text-align: center;">OR</p> <table border="1" style="width: 100%;"> <thead> <tr> <th style="width: 30%;"></th> <th style="width: 35%; text-align: center;">Cold Water</th> <th style="width: 35%; text-align: center;">Hot Water</th> </tr> </thead> <tbody> <tr> <td style="text-align: center;">Sodium</td> <td> <ul style="list-style-type: none"> Reacts violently Heat is evolved. </td> <td> <ul style="list-style-type: none"> React more violently More heat is evolved. (Any one point) </td> </tr> <tr> <td style="text-align: center;">Calcium</td> <td> <ul style="list-style-type: none"> Reacts less violently as compared to sodium. </td> <td> <ul style="list-style-type: none"> Reacts violently with hot water and sticks to surface of metal and floats on surface of water. </td> </tr> <tr> <td style="text-align: center;">Magnesium</td> <td> <ul style="list-style-type: none"> Does not react with cold water </td> <td> <ul style="list-style-type: none"> React with hot water and floats on surface of water. </td> </tr> </tbody> </table>	Galvanisation	Alloying	1. Coating a layer of zinc metal on the metal.	1. Mixing of a metal with metal or non-metal.	2. Not a homogeneous mixture.	2. Homogeneous mixture.	3. No change in physical properties of metals takes place.	3. Change in physical properties of metals takes place.	4. The process is an outcome of the reactivity of metals.	4. Reactivity of metals do not play any role in it.	5. Prevents rusting only.	5. Some alloys may prevent rusting and also used for other advantages.		Cold Water	Hot Water	Sodium	<ul style="list-style-type: none"> Reacts violently Heat is evolved. 	<ul style="list-style-type: none"> React more violently More heat is evolved. (Any one point) 	Calcium	<ul style="list-style-type: none"> Reacts less violently as compared to sodium. 	<ul style="list-style-type: none"> Reacts violently with hot water and sticks to surface of metal and floats on surface of water. 	Magnesium	<ul style="list-style-type: none"> Does not react with cold water 	<ul style="list-style-type: none"> React with hot water and floats on surface of water. 	<p style="text-align: center;">1×3</p> <p style="text-align: center;">1×3</p>	3
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17.	<ul style="list-style-type: none"> Carbon atom, the first member of group 14 has the smallest size in the group and highest inter-atomic force of attraction. It has four valence electrons and requires four more electrons to attain stable configuration. Due to its small size, nucleus of carbon is able to hold the shared pairs of electrons strongly. The bonds formed by other elements of the same group are weaker due to bigger size of their atoms. <p style="text-align: right;">(Any Three)</p>	1×3	3																								

18.	<ul style="list-style-type: none"> A cheetah on seeing a prey generates a <u>nerve impulse</u> which reaches the muscles and the muscle fibre moves. The muscle cell will then move by <u>changing their shape</u> so that muscle cells shorten. Muscle cells have <u>special proteins</u> that change both shape and their arrangement in the cell in response to nervous electrical impulses. When this happens <u>new arrangements</u> of these proteins give the muscle cells a shorter form. 	1 ½ 1 ½	3
19.	<p>The movement of the growth of the roots downwards and the shoots upwards under the stimuli of gravity is called geotropism./ The movement of the part of the plant towards or away from the stimulus gravity.</p>  <p style="text-align: right;">Diagram Labelling</p>	1 1 ½+ ½	3
20.	<ul style="list-style-type: none"> Evolution – Gradual change in living organisms with time since the beginning of life resulting in the formation of a new species/ Evolution is simply the generation of diversity and the shaping of diversity by environmental selection. Evolution cannot be equated with progress because more and more complex body designs have emerged and evolved over time but this does not mean that older designs are inefficient. Foreg. Bacteria are simpler organisms but some inhabit the most inhospitable habitats like hot springs, deep sea, thermal vents and the ice in Antartica. <p style="text-align: center;">OR</p> <p>Examples of feathers :</p> <ul style="list-style-type: none"> Feathers can start out as providing insulation in cold weather but later they might become useful for flight. Some dinosaurs have feathers but they could not fly. Birds later adapted the feathers for flight . This shows that birds are closely related to reptiles. 	1 2 ½ 1 1 ½	3
21.	<p>(a) Behind the mirror (b) Magnified (c) Virtual and erect</p> <p>Labelled ray diagram</p> 	½ ½ ½ 1½	3

22.	<ul style="list-style-type: none"> • $n_{xy} = \frac{2}{3} \quad \therefore \quad n_{yx} = \frac{3}{2}$ • $n_{yz} = \frac{4}{3} \quad \therefore \quad n_{zy} = \frac{3}{4}$ $n_{zx} = n_{zy} \times n_{yx}$ $\therefore \quad n_{zx} = \frac{3}{4} \times \frac{3}{2} = \frac{9}{8}$ <ul style="list-style-type: none"> • $n_{yx} = \frac{v_x}{v_y}$ $\frac{3}{2} = \frac{3 \times 10^8}{v_y}$ $v_y = \frac{3 \times 10^8 \times 2}{3} = 2 \times 10^8 \text{ m/s}$	<p>1/2</p> <p>1/2</p> <p>1/2</p> <p>1/2</p> <p>1/2</p> <p>1/2</p>	3
23.	<p>(a) Presbyopia</p> <p>(b) Gradual weakening of the ciliary muscles of the eye/ diminishing flexibility of the eye lens.</p> <p>(c) Bifocal lens</p> 	<p>1/2</p> <p>1</p> <p>1/2</p> <p>1</p>	3
24.	<p>By placing second (identical) prism in an inverted position with respect to the first prism.</p>  <p style="text-align: right;">Diagram Labelling</p>	<p>1</p> <p>1 1/2</p> <p>1/2</p>	3
SECTION C			
25.	<p>Olfactory indicator</p> <p>(a) Colourless and Odourless gas is evolved with bubbles .</p> $\text{Zinc} + \text{Acid} \rightarrow \text{Zinc Salt} + \text{H}_2 \uparrow$ <p>(or by using any example of acid e.g. HCl/ H₂SO₄)</p> <p>(b) Brisk effervescence/ colourless and odourless gas is evolved.</p> <p>Sodium carbonate + Acid \longrightarrow Sodium salt of Acid + Water + Carbon dioxide\uparrow</p> <p>(or by using any example of acid like HCl/ H₂SO₄)</p>	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p>	5

	OR		
	<ul style="list-style-type: none"> Water of crystallization is the fixed number of water molecules present in one formula unit of a salt. Examples $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ $\text{Na}_2\text{CO}_3 \cdot 10 \text{H}_2\text{O}$ (or Any other) Heat a few crystals of hydrated copper sulphate(bluecolour) in a dry boiling tube. Water droplets are seen in the boiling tube. Colour : The colour of copper sulphate changes to white . State : The blue crystal changes to white powder. 	1 $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ $\frac{1}{2}$ 1 1	
26.	(a)(i) Ductility / Malleability / Lusture (Any two) (ii) Silver , Copper (iii) Gallium, Caesium (b)	$\frac{1}{2}$, $\frac{1}{2}$ $\frac{1}{2}$, $\frac{1}{2}$ $\frac{1}{2}$, $\frac{1}{2}$ 2	5
27.	(a)For providing energy for various metabolic processes / Formation of new cells / Repair of damaged or worn out cells & tissues / Developing resistance against diseases. (Any Two) (b) Peristaltic movement / Peristalsis/ Rhythmic contraction and relaxation of the muscles in the lining of alimentary canal. (c) Herbivores eat plant matter which is rich in cellulose and takes longer time to digest and hence longer small intestine. (d) The inner lining of the stomach will not be protected from the action of the acid /HCl	1+1 1 1 1	5
28.	<ul style="list-style-type: none"> Process of fertilization : Pollen tube is formed from the pollen grain. Fusion of male germ cell with female germ cell to form zygote. 	Labelling Diagram 1 1 1	

	<ul style="list-style-type: none"> • Ovary- Changes into fruit • Ovule- Changes into Seed <p style="text-align: center;">OR</p> <p>(a) The period during adolescence is called puberty/age of males and females at which reproductive organs become functional.</p> <p>(b) (i) Testes – Production of sperms /Secretion of male sex hormone testosterone.</p> <p>(ii) Seminal vesicle – secretes a fluid which makes the transport of sperms easier/ the fluid secreted gives nutrition to sperms.</p> <p>(iii) Vas deferens- carries the sperms to the seminal vesicle.</p> <p>(iv) Urethra- forms a common passage for both the sperms and urine.</p> <p style="text-align: center;">(Any one)</p> <p>(c) Because the sperm formation requires a lower temperature than the normal body temperature.</p> <p>(d) With the help of a long tail.</p>	<p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>5</p> <p>1</p> <p>$\frac{1}{2} \times 4$</p> <p>1</p> <p>1</p>	
29.	 <p>(a) Three resistors are connected in parallel hence voltage across each is same i.e. 6V.</p> $I_1 = \frac{V}{R_1} = \frac{6}{10} = 0.6 \text{ A}$ $I_2 = \frac{V}{R_2} = \frac{6}{20} = 0.3 \text{ A}$ $I_3 = \frac{V}{R_3} = \frac{6}{30} = 0.2 \text{ A}$ <p>b) $I = I_1 + I_2 + I_3 = 1.1 \text{ A}$</p> <p>c) $R_{\text{eff}} = \frac{V}{I}$</p> $= \frac{6}{1.1} = 5.4\Omega$ <p style="text-align: center;">OR</p> <p>$R_1 = R_2 = 15\Omega \quad V = 6V$</p> <p>i) In series :</p>	<p>2</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p> <p>$\frac{1}{2}$</p>	

	$R_s = R_1 + R_2 = 15\Omega + 15\Omega = 30\Omega$ $I = \frac{V}{R_3} = \frac{6V}{30\Omega} = 0.2 \text{ A}$ $\therefore P_1 = VI = 6V \times 0.2 \text{ A} = 1.2 \text{ W}$ <p>ii) In parallel</p> $R_p = \frac{R_1 \times R_2}{R_1 + R_2} = \frac{15 \times 15}{15 + 15} = \frac{225}{30} = 7.5\Omega$ $I = \frac{V}{R_p} = \frac{6V}{7.5\Omega} = 0.8 \text{ A}$ $P_2 = VI = 6V \times 0.8 \text{ A} = 4.8\text{W}$ $\text{Ratio of power} = \frac{P_1}{P_2} = \frac{1.2\text{W}}{4.8\text{W}} = \frac{1}{4}$ $\therefore P_1 : P_2 = 1 : 4$	$\frac{1}{2}$ $\frac{1}{2}$ 1 $\frac{1}{2}$ $\frac{1}{2}$ 1 $\frac{1}{2}$ $\frac{1}{2}$	5
30	<p>(a)Flemings' Left hand rule: Stretch the thumb, forefinger and middle finger of your left hand such that they are mutually perpendicular. If the forefinger points in the direction of magnetic field, middle finger in the direction of current , then the thumb will point in the direction of motion or force acting on the conductor.</p> <p>(b) Three charactersticfeatures :</p> <ul style="list-style-type: none"> • Reverses direction periodically. • Frequency of 50 Hz. • Potential difference between live wire and netural wire is about 220V. <p>(c)</p> <ul style="list-style-type: none"> • Fuse is a safety device used in a circuit (or appliance) to prevent damage due to overloading/ short circuiting. • It protects the circuit (or appliance) by stopping the flow of any unduly high electric current / If current larger than the specified valueflows through the circuit , due to Joule's heating effect the fuse wire melts and breaks the circuit. <p>(d) It provides a low resistance conducting path for the current and protects the user from electric shock due to leakage of current.</p>	$1\frac{1}{2}$ $\frac{1}{2} \times 3$ $\frac{1}{2}$ $\frac{1}{2}$ 1	5