Primary Valency	Ionisable and equal to the oxidation state of the metal in the coordination compound It is satisfied by negative ions
Secondary Valency	Directional and non- ionisable and equal to the coordination number It is satisfied by neutral or negative ligands
Double Salt	Dissociate into simple ions completely when dissolved in water E.g. Carnallite - $KCl.MgCl_2.6H_2O$, Mohr's salt - $FeSO_4.(NH_4)2SO_4.6H_2O$, Potash alum - $KAl(SO_4)_2.12H_2O$

Unidentate Ligand	It binds to the metal ion through a single donor atom E.g. Cl^-, H_2O or NH_3
Bidentate Ligand	It binds to the central atom through two donor atoms E.g. ethane-1,2-diamine, oxalate
Polydentate Ligand	A single ligand with many donor atoms E.g. EDTA is a hexadentate ligand

Ambidentate Ligand	The ligand, which has two different donor atoms and can bind to the central atom through either of the two donors E.g. NO_2^- and SCN^- ions
Coordination Number	It is the number of ligand donor atoms to which the central metal is directly bonded in a complex E.g. In $[PtCl_6]^{2-}$, the coordination number of Pt is 6 In $[Fe(C_2O_4)_3]^{3-}$, the coordination number of Fe is 6 because oxalate is a bidentate ligand.
Homoleptic Complexes	It is a complex in which a metal is bound to only one kind of donor groups E.g. $[Co(NH_3)_6]^{3+}$, $[Ni(CO)_4]$, $[PtCl_6]^{2-}$

Heteroleptic Complexes	It is a complex in which a metal is bound to more than one kind of donor groups E.g. $[Co(NH_3)_4Cl_2]^+$