

## Chemistry Practical Class 10 Arranging Zn, Fe, Cu and Al metals in the decreasing order of reactivity Viva Questions with Answers

**Q1.** Name any two metals that are more reactive than iron.

**Answer:** Aluminium and zinc are more reactive than iron.

**Q2.** What happens when copper is added to ferrous sulphate solution?

**Answer:** No reaction will occur as copper is less reactive than iron.

**Q3.** What is the colour of copper sulphate solution?

**Answer:** Copper sulphate solution is blue coloured.

**Q4.** What is the colour of the aluminium sulphate solution?

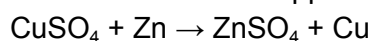
**Answer:** Aluminium sulphate solution is colourless.

**Q5.** What is the colour of the ferrous sulphate solution?

**Answer:** Ferrous sulphate solution is green coloured.

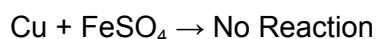
**Q6.** Why did the cuprous sulphate solution colour change when zinc metal was dipped?

**Answer:** Cuprous sulphate solution colour changed when zinc metal was dipped into it because zinc is more reactive than copper. Zinc displaces copper from copper sulphate and forms zinc sulphate.



**Q7.** What will you observe when copper is added to the ferrous sulphate solution?

**Answer:** No reaction will occur when copper is added to the ferrous sulphate solution.



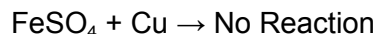
**Q8.** Which are the most and the least reactive metal in the above experiment?

**Answer:** Aluminium is the most reactive, while copper is the least reactive metal in the above experiment.

**Q9.** Why can we safely preserve ferrous sulphate in a copper vessel, whereas we can not preserve the same in a zinc vessel?

**Answer:** We can preserve ferrous sulphate in a copper vessel because copper is less reactive than iron. Thus no reaction will take place. But, we can not store ferrous sulphate in a zinc vessel because

zinc is more reactive than iron. Thus, it will displace iron from ferrous sulphate and form zinc sulphate solution.

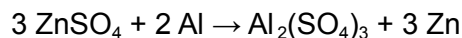


**Q10.** What does the reactivity series of metals indicate?

**Answer:** The reactivity series of metals indicates the reactivity of metals in decreasing order.

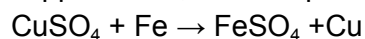
**Q11.** Can we store zinc sulphate in an aluminium container? Give a reason for your answer.

**Answer:** No, we can not store zinc sulphate in an aluminium container because aluminium is more reactive than zinc. Thus, it will displace zinc from zinc sulphate solution and form aluminium sulphate.



**Q12.** Can we store copper sulphate in an iron container? Give a reason for your answer.

**Answer:** No, we can not store copper sulphate in an iron container because iron is more reactive than copper. Thus, it will displace copper from copper sulphate solution and form ferrous sulphate.



**Q13.** Why is the reactivity of different metals different?

**Answer:** The reactivity of different metals is different because of the difference in the tendency to lose electrons.