

Chemistry Practical Class 10 Types of Reaction Double Displacement Reaction - Reaction Between Sodium Sulphate And Barium Chloride Solutions Viva Questions with Answers

Q1. What is Double Displacement Reaction?

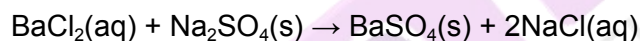
Answer: Double Displacement reaction is a type of chemical reaction in which two reactants exchange ions to form two new compounds.

Q2. Write a general equation for the double displacement reaction.

Answer: $AB + CD \rightarrow AD + BC$ is the general equation for displacement reaction.

Q3. What is the balanced equation for the reaction between barium chloride and sodium sulphate?

Answer: The balanced equation is:



Q4. What is the colour of BaSO_4 precipitate?

Answer: BaSO_4 precipitate is white in colour.

Q5. Is the precipitate formed soluble in HCl?

Answer: No, the precipitate formed is not soluble in HCl.

Q6. Why does the reaction between barium chloride and sodium sulphate reach completion?

Answer: The reaction between barium chloride and sodium sulphate reaches completion because it is irreversible and the white precipitate of barium sulphate is insoluble in aqueous solution.

Q7. What is the atomic mass of barium?

Answer: The atomic mass of barium is 137 g/mol.

Q8. What is the atomic mass of sodium?

Answer: The atomic mass of sodium is 23 g/mol.

Q9. Is the above discussed double displacement reaction spontaneous?

Answer: Yes, the double displacement reaction takes place at room temperature.

Q10. What are the precautions to be taken before performing the experiment?

Answer: Acid and alkalis should be handled carefully and the glassware should be thoroughly washed before performing the experiment.

