

Chemistry Worksheets Class 10 on Chapter 2 Acids, Bases and Salts - Set 2

Q1. Fresh milk has a pH of 6. When milk changes into curd, the pH value will:

- a.) become 7
- b.) become less than 6
- c.) become more than 7
- d.) remains unchanged

Q2. Which of the following is medicine for indigestion?

- a.) sodium hydroxide
- b.) manganese hydroxide
- c.) magnesium hydroxide
- d.) potassium hydroxide

Q3. How many number of water of crystallisation is present in copper sulphate crystals?

- a.) 3
- b.) 5
- c.) 7
- d.) None

Q4. The indicators which turn red in acid solution are:

- a.) turmeric and litmus
- b.) phenolphthalein and methyl orange
- c.) litmus and methyl orange
- d.) phenolphthalein and litmus

Q5. The salt which will give an acidic solution on dissolving in water is:

- a.) KCl
- b.) NH_4Cl
- c.) Na_2CO_3
- d.) CH_3COONa

Q6. What are the observations when quick lime is added to water?

Q7. Which indicator gives pink colour in the basic solution?

Q8. Fill in the blanks.

- a.) The chemical formula of washing soda is ____.
- b.) The chemical formula of sodium carbonate decahydrate is ____.

Q9. Write the formula of the calcium sulphate and identify the acid and base.

Q10. Name the acids present in the following:

- i.) Lemon juice
- ii.) Vinegar
- iii.) Vitamin C tablet
- iv.) Tamarind
- v.) Sour milk

Q11. What happens when carbon dioxide is passed through lime water?

Q12. How does toothpaste prevent tooth decay?

Q13. Explain how the pH change in the lake water can endanger the lives of aquatic animals. What can be done to lessen the danger to the lives of aquatic animals in the lake?

Q14. Describe how sodium hydrogen carbonate is produced on a large scale.

Q15. How would you show that blue copper sulphate crystals contain water of crystallisation?

Q16. What is the 'Chlor-alkali' process and what products are formed during the process?

Q17.

- a.) What happens during a bee sting? What is its remedy?
- b.) What happens during a wasp sting? What is its remedy?
- c.) Name any two chemical materials which can mix with the acidic soil to adjust its pH?

Q18. You are given two solutions A and B. The pH of solution A is 6 and pH of solution B is 8.

- i.) Which solution is acidic and which is basic?
- ii.) Which solution has more H^+ ion concentration?
- iii.) Why is HCl a stronger acid than acetic acid?

Q19. a.) A sample of bleaching powder was kept in an air-tight container. After a month, it lost some of its chlorine content. How will you account for it?

b.) A solution has a pH of 7. Explain how you would

- i.) increase its pH
- ii.) decrease its pH

c.) On exposure to the atmosphere, Glauber's salt loses weight while quicklime gains weight.

Q20. Answer the following questions:

- a.) What is plaster of paris? Write its chemical formula.
- b.) What happens when bleaching powder reacts with dilute sulphuric acid? Give an equation of the reaction involved.
- c.) Give two important uses of washing sods.
- d.) What will be the colour of the litmus in an aqueous solution of ammonium chloride salt?
- e.) What is meant by hydrates and anhydrous salts? Explain with examples.

