
1. Low density of ice compared to water is due to:
 - A. Dipole - induced dipole interactions
 - B. Induced dipole - induced dipole interactions
 - C. Hydrogen bonding interactions
 - D. None of the above

2. Which of the following bonds/forces is the weakest?
 - A. Covalent bond
 - B. Ionic bond
 - C. Hydrogen bond
 - D. London force

3. Intermolecular hydrogen bonding is not present in which of the following pairs of molecules?
 - A. SiH_4 and SiF_4
 - B. $CH_3 - \overset{\overset{O}{||}}{C} - CH_3$ and $CHCl_3$
 - C. $H - \overset{\overset{O}{||}}{C} - OH$ and $CH_3 - \overset{\overset{O}{||}}{C} - OH$
 - D. All have intermolecular H-Bonding

4. Which property is not due to H-bonding?
 - A. High boiling point of water
 - B. High viscosity of glycerol
 - C. Solubility of ammonia in water
 - D. Polar nature of halogen acid

5. The strongest hydrogen bonding is present in:
 - A. HF
 - B. HCl
 - C. HBr
 - D. HI