

1. Low density of ice compared to water is due to:
 - A. Dipole - induced dipole interactions
 - B. Induced dipole - induced dipole interactions
 - C. Hydrogen bonding interactions
 - D. None of the above

2. Which of the following bonds/forces is the weakest?
 - A. Covalent bond
 - B. Ionic bond
 - C. Hydrogen bond
 - D. London force

3. Intermolecular hydrogen bonding is not present in which of the following pairs of molecules?
 - A. SiH_4 and SiF_4
 - B. $\text{CH}_3 - \overset{\text{O}}{\parallel} \text{C} - \text{CH}_3$ and CHCl_3
 - C. $\text{H} - \overset{\text{O}}{\parallel} \text{C} - \text{OH}$ and $\text{CH}_3 - \overset{\text{O}}{\parallel} \text{C} - \text{OH}$
 - D. All have intermolecular H-Bonding

4. Which property is not due to H-bonding?

- A. High boiling point of water
- B. High viscosity of glycerol
- C. Solubility of ammonia in water
- D. Polar nature of halogen acid

5. The strongest hydrogen bonding is present in:

- A. HF
- B. HCl
- C. HBr
- D. HI